

## Atomic Structure

- List and explain all three sub atomic particles of an atom
- Explain how those particles are arranged in an atom
- Draw a sodium Bohr model atom and label the P,N,E, electron cloud, and nucleus

## Periodic Table

- Explain how Mendeleev's periodic table differs from Moseley's
- What is the difference between the atomic number and atomic mass
- Why are groups 1 and 17 the most reactive groups on the periodic table?
- What group is considered the least reactive group on the periodic table. WHY?

## Molecules, compounds, mixtures

- Explain the difference between a compound and a molecule
- What is the difference between a heterogeneous and homogenous mixture. Provide 2 examples of each.
- Is chocolate milk a homogeneous or heterogeneous mixture? Identify the solvent, solute, and solution.

## Element Practice

Element Name	silicon	Arsenic	Magnesium
Atomic symbol			
Atomic #			
Atomic Mass			
Protons			
Neutrons			
Electrons			
Valence Electrons			
G (group)			
P (period)			
Metal, nonmetal, metalloid			

Fill out on  
other page

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## Periodic Table

## Molecules, compounds, mixtures

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